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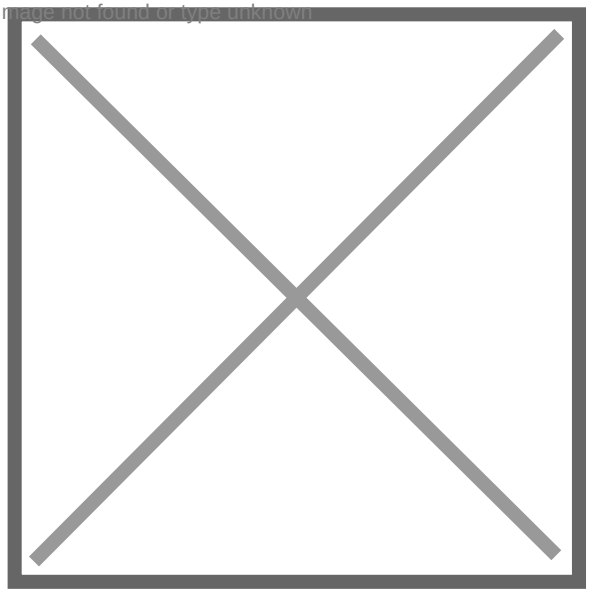
Product Name :

Potassium Permanganate

Product Code :

SCI-560/163

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Science Lab
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Description :

Potassium Permanganate is an analytical reagent grade oxidising agent, supplied ina 200 g pack with an MSDS. Its deep purple solution loses colour as it reacts with reducing agents, so it acts as its own indicator in redox titrations. School and college labs also use the crystals for diffusion and convection demonstrations, oxygen preparation and Baeyer's test for unsaturation.

Potassium permanganate grade and pack

Parameter	Detail
Grade	Analytical reagent (AR)
Pack size	200 g
Safety data	MSDS supplied with the pack
Product code	SCI-560/163
Minimum order	10 units
Dispatch	Usually within 7 to 10 working days
Supplied by	ScienceLab (Jain Scientific Equipments Pvt Ltd), Ambala, India

Practicals that use potassium permanganate

Titration against Mohr's salt. Make up a standard solution of ferrous ammonium sulphate in water containing some dilute sulphuric acid, and pipette a known volume into a conical flask. Run the permanganate in from a burette until one drop leaves a faint pink that stays on swirling. Repeat until the titres agree, then work out the molarity of the permanganate.

Titration against oxalic acid. Pipette standard oxalic acid solution into a flask, add dilute sulphuric acid and warm it. Titrate while the mixture is still hot. The first drops lose their colour slowly, but later ones vanish almost at once because the manganese(II) ions formed speed up the reaction.

Diffusion and convection. Drop one crystal into a glass beaker of still water and watch purple streaks spread with no stirring. For convection, set a crystal against one side of the beaker floor and heat gently under that spot. The coloured water rises above the flame, crosses the top and sinks on the far side.

Baeyer's test. Shake a few drops of dilute alkaline permanganate with the organic sample in a test tube. Loss of the purple colour and a brown precipitate point to a carbon-carbon double or triple bond. Aldehydes and some other easily oxidised compounds also decolourise it, so a positive result is a strong hint rather than proof.

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